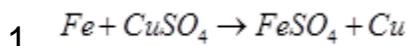


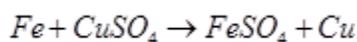
Class 10 Science
Ch-1 Chemical Reactions And Equations
Practice Paper

More Questions

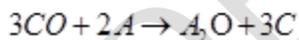
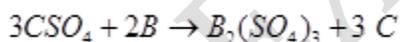


Identify the type of reaction.

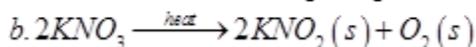
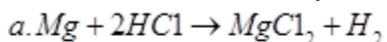
2. What does the symbol (s) used with water indicate?
3. How can we prevent fried food from turns 'Rancid'?
4. Why does lime water turn milky when CO_2 is passed into it?
5. Which gas is evolved when lead nitrate is heated?
6. During electrolysis of water, how can we identify the gas present in each test tube?
7. Give an example of a photolytic reaction which is not a decomposition reaction?



8. In above reaction iron nail becomes brownish in colour and the blue colour of copper sulphate solution fades. Why?
9. Identify the element which is most reactive and the element which is least reactive?



10. Write a chemical equation of a reaction in which a precipitate is formed.
11. Write your observation when Magnesium ribbon is burned in air? Name the powder formed.
12. Which characteristics of a chemical change do you observe when dilute sulphuric acid is added to zinc granules in a conical flask?
13. Write word equation for the following chemical equation:



14. What happens when CO_2 (g) is bubbled through lime water. Write the chemical equation.
15. What happens when a silver spoon is kept immersed in aqueous Conner sulphate solution?
16. Why does Conner not liberate hydrogen on readme with dilute sulphuric acid?
17. Write a chemical equation to show the process of respiration. Mention the type of reaction.
18. Which of the following reactions show evolution of gas.
- a. $2\text{AgCl} \rightarrow 2\text{Ag} + \text{Cl}_2$
 - b. $\text{Pb} + \text{CuCl}_2 \rightarrow \text{PbCl}_2 + \text{Cu}$
 - c. $\text{CuO} + \text{H}_2 \rightarrow \text{Cu} + \text{H}_2\text{O}$
 - d. $\text{ZnO} + \text{C} \rightarrow \text{Zn} + \text{CO}$
19. Name 2 metals which at tarnished. Why does this happen?
20. Why is corrosion harmful?
21. Mention three situations in daily life where a chemical change occurs.
22. Balance the following chemical equations.
- a. $\text{H}_2\text{SO}_4 + \text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$
 - b. $\text{NaCl} + \text{AgNO}_3 \rightarrow \text{AgCl} + \text{NaNO}_3$
 - c. $\text{CH}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
23. Write chemical equations for the following word equations:
- a. Hydrogen + Chlorine \rightarrow Hydrogen Chloride
 - b. Sodium + Water \rightarrow Sodium Hydroxide + Hydrogen
 - c. Zinc Oxide + Carbon \rightarrow Zinc + Carbon Monoxide
24. What do you mean by endothermic and exothermic reactions? Give examples.
25. What happens when potassium iodide solution is added to lead nitrate solution? Give equation of reaction and mention the type of reaction involved?
26. How can we make a chemical equation more informative?
27. Write one chemical equation to show:
28. a. Combination reaction b. Decomposition reaction c. Double Displacement Reaction
29. Write short notes on:

- a. Corrosion
 - b. Rancidity
30. A substance X when mixed with water is used for white washing. The substance X is also formed when a substance Y decomposes.
- a. Identify X and Y and write their formula.
31. Define oxidation and reduction. Give an example of a Redox reaction.

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